

Important improvements occur as regards methods for the detection of phosphorus and of other non-metallic poisons.

Abundant literature references give increased value to the entire work. The older system of equivalent weights, retained by Fresenius even in the latest German edition, has been wisely suppressed as serving no good purpose.

The student of Prof. Wells' book cannot fail to be impressed by its accuracy of statement and the painstaking care with which the directions given him are formulated.

The work of the translator has been performed in a thorough manner, and it may be said that the new English version will be welcomed by chemists everywhere. F. C. PHILLIPS.

THE PHARMACIST AT WORK. BY WILLIAM C. ALPERS. Philadelphia: J. B. Lippincott Co. 1898. ix + 326. Price, \$1.50.

The usual review methods of scientific works are hardly applicable to this book. The story, in its outward features at least, is a story of the past and appeals with especial force to men who look backward in pharmacy. The preceptor and apprentice are before us again; throughout twenty-four chapters and 326 pages, we are captivated by a pithy dialogue wherein the author not only treats of pharmaceutical manipulation and the chemical and physical principles underlying them, but also comments on many possibilities with which a pharmacist's life is confronted, portraying with much refinement not a few of the frailties of human nature. From beginning to end, every page bears testimony to the author's love for his art and the earnestness of the moral and ethical principles he advocates.

Mr. Alpers, a typical German pharmacist, who came to America just after the Franco-German war, and who has thorough command of his subject, takes the opportunity to make some timely suggestions to modern druggists. He points out (p. 41) the great unknown in alkaloidal and Galenical chemistry as applied to pharmacy, and it is to be hoped that his earnest words will arouse many pharmacists who have the time but waste it by indifference to scientific thought. He informs the reader (p. 56) that tinctures are superior to tablets as medicine carriers, but he realizes also that tablets are so easily counted into a box without involving the "bother" of laboratory work, as to render it possible that better preparations will suffer.

While considering plasters, pills, emulsions, decoctions, syrups, and other classes of preparations, the author instructs and interests not only the clerks whom he addresses, but any pharmacist who reads after him and wishes to refresh his mind or to learn anew. Scarcely a subject of value in drug-store manipulation but is well treated, and the errors detected in the book are trivial or typographical, such as the word carbonate (p. 197) instead of carbamate. The reviewer picked up the book in the expectation of a task in the "twice told tale;" he read it through with unalloyed interest, forgetting the task. It is a book for pharmacists to read, to think over, to read again; it is a book for the young druggist without an instructor, as well as to refresh the mind of the teaching pharmacist who is an instructor professionally. Especially should it be read by those who are contented with the belief that the art of pharmacy has no higher standard than the counting out of factory-made pills and tablets.

JOHN URI LLOYD.

A DICTIONARY OF CHEMICAL SOLUBILITIES, INORGANIC. BY A. M. COMEY. New York: The Macmillan Co. 1896. xx + 515 pp. Price, \$5.00.

This new Dictionary of Solubilities is limited to substances not containing carbon, exception being made in the case of the carbonates, cyanides, carbon monoxide and disulphide, and a few others. While a complete compilation of data for all carbon compounds would have delayed its appearance, it is to be regretted that the author did not include some of these which are of importance to inorganic chemists. A few pages devoted to the oxalates, acetates, citrates, tartrates, the commoner alcohols and a few other organic compounds, would have added much to the value of the work without materially increasing its size; we are at a loss to understand why carbonyl chloride, nickel carbonyl, or an aromatic derivative like potassium carbonyl should be preferred to these. The plan of placing oxy- and sulpho-salts under the acid and other metallic compounds under the metal, has the precedent of use by some dictionaries of chemistry, but we think that classification under the negative radical in all possible cases would have been better as more consistent and as facilitating reference and comparison. There is